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- For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.*



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(54) Title: PROCESS FOR PREPARATING ALKALI- AND HEAT-STABLE SUGAR ALCOHOL COMPOSITIONS AND A SORBITOL COMPOSITION

(57) Abstract: The invention relates first of all to a process for preparing alkali- and heat-stable sugar alcohol compositions which exhibits an optical density lower than or equal to 0,100 in an S-test, in which a sugar alcohol composition is treated with a strong base anion exchange resin in the hydroxide form, at a temperature between 30 °C and 100 °C. Second of all, the invention relates to a sorbitol composition

**PROCESS FOR PREPARING ALKALI- AND HEAT-STABLE SUGAR
ALCOHOL COMPOSITIONS AND A SORBITOL COMPOSITION**

5 The invention relates on the one hand to a process for preparing alkali- and heat-
stable sugar alcohol compositions which exhibits an optical density lower than or
equal to 0,100 in an S-test. On the other hand, the invention relates to sorbitol
compositions.

10 Alkali- and heat-stability of sugar-alcohols is important in all those applications
where colour formation under these conditions must be prohibited. This is the case,
e.g. where polyol compositions are used as humectants in tooth-pastes containing
alkaline abrasives, as building blocks of polyether polyols, or as starters for
preparing sorbitan esters. Colouring of end-products containing these polyols is
15 often due to the presence of colour-forming precursors, including residual reducing
sugars, in the sugar-alcohol compositions used.

This problem is quite well known and a number of solutions have already been
proposed to improve alkali- and heat-stability of such polyol compositions.

20 In JP 63079844, a method is described in which an aqueous sugar alcohol solution is
adjusted to a pH-value of between 8 and 13, followed by a (discontinuous) heating
step at temperatures varying between 90°C and 220°C. The resulting product is then
purified by passing the polyol solution through a strongly acidic cation exchange
resin, a strong base anion exchange resin and a mixed bed resin.

25 In EP 0 711 743, a similar process is disclosed in which the polyol composition is
first stabilised by means of an oxidation, a fermentation or a caramelisation step,
followed by a purification of the solution. The purification step is comparable to the
one disclosed in JP 63079844.

In EP 1 095 925, a purification process is disclosed comprising a first treatment on a strong acid cationic exchange resin at a temperature below 50°C, preferably below 40°C, followed by a treatment on a strong basic anionic resin and a mixed bed resin.

5 Methods for removing aldehydes and other reactive impurities, and/or stabilising colour in glycerol and glycol aqueous solutions have been discussed in FR 1 546 472 and US 6,187,973 respectively. In both cases, a strong base anion exchange resin in the hydroxide form was converted into the bisulphite form, in order to treat the aqueous polyol solutions.

10

The major disadvantage of the above-cited processes resides in their complexity. Indeed, first a chemical stabilisation step is needed, followed by a quite complicated purification step. During this stabilisation step, high pH-values are used. In addition, it is necessary to use quite high temperatures (> 90°C) and long reaction
15 times to obtain the necessary stabilisation.(see JP63079844 and EP711743). This results in a rather important chemicals consumption during the chemical reaction, and later on, for the regeneration of the different ion exchange resins.

In addition, two separate steps are needed to arrive at the desired result. The
20 equipment needed, therefore comprises a reactor to perform the stabilisation step, and at least two ion exchange resin batteries to perform the purification step.

In the case of the bisulphite-type resins, the use thereof proves to be inefficient when higher polyols such as pentitols, hexitols and/or hydrogenated starch hydrolysates are
25 treated in the expectation of obtaining alkali- and heat stable products.

The purpose of the invention is to provide a simple process for preparing alkali- and heat-stable sugar alcohol compositions which exhibit an optical density lower than or equal to 0,100 in an S-test. This process results in a reduced consumption of
30 chemicals and provides a combined one step alkaline stabilisation and

decolourisation process. This one step process can be operated in a continuous matter.

This object is obtained by providing a process for preparing alkali- and heat-stable sugar alcohol compositions which exhibits an optical density lower than or equal to 0,100 in an S-test, in which a sugar alcohol composition is treated with a strong base anion exchange resin in the hydroxide form, at a temperature between 30 °C and 100 °C.

10 In a preferred process according to the invention, the sugar alcohol composition is fed to a column-system containing a strong base anion exchange resin in the hydroxide form with a volume throughput of ≤ 6 bed volumes (BV)/hour.

"Bed volume" is hereby defined as the total volume of resin used during the stabilisation step, be it in a one column- or in a multiple column-system.

When a multiple column-system is used, at least part of the columns of the system is used in a regeneration mode, while the remaining columns are used in a service mode, comprising the steps of stabilisation and simultaneous decolourisation.

20 In a more preferred process according to the invention, the volume throughput is between 0,1 and 1 BV/hour.

Most preferably, the volume throughput is between 0,2 and 0,8 BV/hour.

25 Before treatment with the strong base anion exchange resin, said sugar alcohol composition has preferably a conductivity value less than 100 $\mu\text{S}/\text{cm}$, more preferably 50 $\mu\text{S}/\text{cm}$.

30 The strong base anion exchange resin preferably belongs to one of the categories:
- the thermally stable-type category;

- the styrenic type I, type II or type III; or
- the acrylic resin type.

When using a styrenic type I or type III, or an acrylic type resin, a column
5 temperature is preferably used between 45 °C and 70 °C.

When using a styrenic type II resin, a column temperature is preferably used less
than 45 °C.

10 When using a thermally stable resin, a column temperature is preferably used which
is more than 75 °C.

On the one hand, said sugar alcohol composition can be prepared by hydrogenating a
starch hydrolysate, obtained from an acid conversion, a combined acid-enzymatic
15 conversion or a multiple enzyme conversion of starch.

On the other hand, said sugar alcohol composition can be prepared by hydrogenating
reducing sugars belonging to the categories of keto- or aldopentoses, keto- or
aldohexoses, disaccharides or non-starch oligosaccharide mixtures.

20

In a preferred process according to the invention, said sugar alcohol composition has
a pH-value between 8,5 and 9,5 when sorting from the strong base anion exchange
resin.

25 The purpose of the invention is furthermore to provide a sorbitol composition which
has a great alkali- and heat stability.

This purpose is obtained by providing a sorbitol composition containing at least 95%
sorbitol on dry substance and exhibiting an optical density lower than 0,02, more
30 preferably containing at least 99 % sorbitol on dry substance and exhibiting an
optical density of lower than 0,01.

These particular and unexpected sorbitol compositions can for instance be the result of the process of the invention. This doesn't exclude the fact that also other processes can be used for obtaining sorbitol compositions with the characteristics as mentioned above.

To prepare alkali- and heat-stable sugar alcohol compositions which exhibits an optical density lower than or equal to 0,100 in an S-test, which is within the terms of EP 0 711 743, a one step process is used in which all the colour precursors are eliminated, and the resulting sugar alcohol syrup become colourless and alkali- and heat-stable. Thereby surprisingly low reaction temperatures are used, i.e. between 30 and 100°C. These temperatures are sufficient to provide the desired result, as expressed by the S-value, when a strong base anion exchange resin in the hydroxyl form is used to treat the sugar alcohol compositions.

The S-test relies on a spectrophotometric measurement applied to the products to be tested.

As used herein, the "S-test" is the following test procedure:

- the aqueous polyol syrup to be tested is brought to a solids content of 40 % by weight, if needed by concentration or by aqueous dilution;
- to 5 ml of this solution are added 500 mf of sodium hydrogenecarbonate of ultrapure quality, sold, for example, under the name of RP NormapurTM, analytical grade, by the company Prolabo, 65 Bd Richard Lenoir, Paris, France, and 250 mg of an aqueous solution containing 20 % of ammonia;
- the whole is mixed and heated for 2 hours on a steam bath at 100 °C without stirring being applied;
- the solution is brought to 20 °C and the optical density of the solution thus obtained is measured at a wavelength of 420 nm by virtue of a spectrophotometer such as that marketed by Perkin-Elmer under the trademark Lambda 5 UV/VIS Spectrophotometer.

In a process according to the invention, a single or multiple column-system is used, containing a strong base anion exchange resin in the hydroxide form with a volume throughput of ≤ 6 bed volumes (BV)/hour.

5

An advantageous continuous multiple column-system for this invention is known as an ISEP- or as a CSEP-configuration. Thereby, a part of the columns in the system is used in the regeneration mode, while the remaining columns are used for the stabilisation and simultaneous decolourisation of the substrate, i.e. a service mode.

10 Thereby, a column operating in the service mode will become deactivated, after that a certain quantity of substrate is processed by that column. This deactivation can be observed by monitoring the pH of the syrup leaving that column. This "exhausted" column is then switched to the regeneration mode and replaced by a regenerated column. The volume throughput is preferably between 0,1 and 1 BV/hour, more
15 preferably between 0,2 and 0,8 BV/h.

The strong base anion exchange resins used in this invention belong to the styrenic type I, type II, or type III categories, to the acrylic resin type, and to the thermally stable-type categories.

20 The styrenic type I-category comprises resin types such as Amberlite IRA404, FPA90 and Amberjet 4400, Dowex Marathon 11 and Lewatit M500.

A typical representative of the type II-resins is the styrenic resin type Dowex 22.

A typical representative of the type III-resins is the styrenic resin type Purolite A555.

The acrylic resin type categorie is represented e.g. by Amberlite IRA458 and
25 Amberlite FPA98.

In the category of the thermally stable strong base anion resins, Diaion TSA1200 is a typical example.

When using styrenic type I or type III resins, or acrylic type resins, column
30 temperature is preferably between 45°C and 70°C.

When using styrenic type II resins, column temperature is preferably $< 45^\circ\text{C}$.

When using thermally stable resins, column temperature is preferably $>75^{\circ}\text{C}$.

The sugar alcohol composition is obtained via the hydrogenation of reducing sugar compositions. Typical reducing sugar compositions are starch hydrolysates, but also
5 other reducing sugars including keto- and aldopentoses, keto- and aldohexoses, disaccharides (e.g. lactose, maltose, isomaltose, isomaltulose) and non-starch oligosaccharide mixtures are covered by this term. The term starch hydrolysates refers to those compositions obtained via an acid conversion, a combined acid-enzymatic conversion, or a multiple enzyme conversion of starch, including chemical
10 or enzymatic isomerisation. Typical representatives here are maltodextrins, "standard" glucose syrups, maltose syrups, high DE conversion syrups such as 96DE and 99DE glucose syrups, isoglucoses and crystallisation mother liquors.

The starch may be of cereal, tuber root or leguminous origin.

15 The sugar alcohol substrate needing a stabilisation treatment, preferably has a conductivity value $< 100 \text{ S/cm}$, more preferable $< 50 \mu\text{S/cm}$, before treatment with the strong base anion exchange resin. Such sugar alcohol compositions can be obtained when using a noble metal hydrogenation catalyst, for example Ru, Pt or Pd; or by first removing dissolved residual metal ions from the substrate, in those cases
20 where transition metals such as for example Ni, Co, Cu or Fe are used as the catalysts.

The processed, colour-stable sugar alcohols thus obtained typically have a pH-value varying between 8,5 and 9,5 when sorting from the strong base anion exchange resin.
25 These sugar alcohol compositions can then be used as such or further processed by means of a mixed bed resin or a weak acid cation exchange resin, thereby providing a syrup having a $\text{pH} = 4 - 7$, preferably $\text{pH} = 5 - 6,5$.

Sorbitol compositions with great alkali- and heat stability can be obtained by the
30 process of the invention as described above. This doesn't exclude the fact that also

other processes can be used for obtaining such particular and unexpected sorbitol compositions.

The sorbitol composition comprise at least 95% sorbitol on dry substance and exhibits an optical density lower than 0,02.

In a more preferred embodiment of the invention, the sorbitol composition comprise at least 99 % sorbitol on dry substance and exhibits an optical density of lower than 0,01.

In order to illustrate this invention, a number of examples are provided hereunder.

EXAMPLES:

Comparative example 1:

A sorbitol syrup containing 96% sorbitol and showing an S-value of 1,10 before stabilisation is submitted to an alkaline heat treatment. Thereby the pH is brought to pH = 11,2 with 1N NaOH and the syrup is heated for 2 hours at different temperatures, followed by a refining step. After refining the stability of the thus processed syrup is measured by means of the S-test. Thereby it is clearly illustrated that sufficiently high temperatures are needed to obtain sufficient stabilisation within a reasonable time. The effect of the stabilisation conditions is clearly illustrated in the following table.

Stabilisation conditions	S-value
Unstabilised	1,10
2h at 80°C and pH=11,2	0,35(is>0,1)
2h at 100°C at pH=11,2	0,021

Comparative example 2:

In this example it is shown that the use of strong base anionic exchange resins, in the sulphite form, is not suitable to obtain alkali- and heat-stable polyols, comparable
5 with the polyols disclosed in EP 0 711 743.

A strong base anion exchange resin in the hydroxide form was thereby converted to the bisulphite form by passing a 2 mol/l solution of NaHSO₃ through a bed of the resin until the bisulphite inlet and outlet concentrations were essentially the same. This resin was then washed with 10 volumes of water per volume of resin.

10 A unrefined sorbitol solution, having a S-value of 0,85 (as determined by the S-test of EP711743) was then passed through the bisulphite-treated resin at ambient temperature.

The treated solutions did not fulfil the requirements set forward by EP 0 711 743, in order to claim alkali/heat stability for these polyols, as illustrated in table 1.

15

Table 1: Treatment of sorbitol by anion exchange resin in bisulphite form:

Sorbitol(99,5%)	S-value
Untreated	0,85
Dowex 22 (1)	0,72
Lewatit M500 (1)	0,64

(1): Dowex 22(type II) and Lewatit M500 (type I) are both strong base anion
20 exchange resins, transformed in the bisulphite form as described above.

Example 1:

25 Four different columns are filled with 100 ml of a different strong base anionic exchange resin in the hydroxide form.

The resins used are: Amberlite FPA 90, Amberlite IRA 458, Amberjet 4400 and Purolite A555.

The operating temperature for the columns containing Amberlite FPA90 and Amberjet 4400 is 60°C. The operating temperature for the columns containing
 5 Amberlite IRA458 and Purolite A555 is 55°C.

The unrefined sorbitol syrup used is obtained by the hydrogenation of a 96 DE-glucose syrup and has a concentration of about 50% d.s.. The S-value of the syrup is 1,11.

The syrup is passed through the columns at 0,33 BV/hour. The S-value of the
 10 processed substrate is determined in the mixture of the collected volumes of the first 25 bed volumes.

Table 2: Sorbitol syrup stabilisation via strong base anionic resin treatment as expressed by their S-value:

15

Sorbitol syrup substrate	S-value
untreated	1,11
Amberlite FPA 90	0,042
Amberlite IRA 458	0,016
Amberjet 4400	0,016
Purolite A555	0,014

Example 2:

20 The same procedure as in example 1 was followed. The substrate treated now is sorbitol syrup, obtained by the hydrogenation of a D99 dextrose syrup. The sorbitol syrup has a 99,4% sorbitol content and an S-value of 0,87 before stabilisation.

This syrup is again passed through the columns at 0,33 BV/hour. The processed sorbitol syrup is collected. A mixed sample corresponding to the mixture of the first
 25 30 bed volumes is submitted to the S-test.

Table 3: D99 Sorbitol syrup stabilisation via strong base anionic resin treatment as expressed by their S-value:

D99 Sorbitol syrup substrate	S-value
untreated	0,87
Amberlite FPA 90	0,007
Lewatit M500	0,006
Purolite A555	0,007

5

Example 3:

The same procedure as in example 1 was followed. The substrate treated now is a maltitol syrup, obtained by the hydrogenation of a high maltose syrup containing about 65 % maltose, 8 % glucose and 20 % maltotriose. The maltitol syrup has an S-value of 2 before refining.

This syrup is again passed through the columns at 0,33 BV/hour. The processed maltitol syrup is collected. A mixed sample corresponding to the mixture of the first 25 bed volumes is submitted to the S-test.

20

25

Table 4: Maltitol syrup stabilisation via strong base anionic resin treatment as expressed by their S-value:

Maltitol syrup substrate	S-value
Untreated	2,0
Amberlite FPA 90	0,065
Amberlite IRA 458	0,08
Amberjet 4400	0,11
Purolite A555	0,09

- 5 This example shows that the substrate composition, the original S-value of the substrate, as well as the ion exchange resin used, may affect the refining capacity.

Example 4:

10

One column is filled with 100 ml of a strong base thermally stable anion exchange resin (Diaion TSA1200). The substrate treated is a maltitol syrup, obtained by the hydrogenation of a high maltose syrup containing about 65% maltose, 8% glucose and 20% maltotriose. The maltitol syrup has an S-value of 2,2 before refining.

- 15 The syrup (50% d.s.) is passed, at 90°C, through the column at a speed of 2BV/hour. The evolution of the S-value as a function of the number of bed volumes processed is given in the next table.

BV treated	S-value
5	0,012
10	0,013
20	0,022
30	0,028
40	0,037

Example 5:

One column is filled with 100 ml of a strong base styrenic type II anion exchange resin (Dowex 22). The substrate treated is a sorbitol syrup (96 % sorbitol), obtained
5 by the hydrogenation of a 96DE glucose syrup. The sorbitol syrup has an S-value of 0,81 before stabilisation.

The syrup (50 % d.s.) is passed, at 40°C, through the column at a speed of 0,2 BV/hour. After 24 hour the S-value of the collected volume was determined. The
10 value of the collect volume was $S = 0,036$.

Example 6:

In this example a multiple column system is used to refine a maltitol syrup having a
15 S-value of 2,2 before treatment. The maltitol syrup is obtained by the hydrogenation of a high maltose syrup containing about 65 % maltose, 8 % sorbitol and 20 % maltotriose.

Four columns are each filled with 100 ml Amberlite IRA458 resin in the hydroxyl form. Three out of four columns are connected in series. The maltitol syrup is
20 passed through these three columns at a speed of 100 ml/h. (0,33 BV/h - 300 ml resin in use). The temperature of the columns is 55°C. The treated syrup is collected in fractions of 1 liter, and of each such fraction the S-value is determined. In a spot sample, at the outlet of the first column, the pH is measured on the moment that a 1 liter fraction is collected at the outlet of the third (last) column of the stabilisation
25 cycle. When the pH shows a significant drop compared to the pH of the previous sampling, then:

- this first column (n°1) is taken out of service and is regenerated
- the second column (n°2) now becomes the first column, and receives the substrate to be treated,
- 30 - the third column (n°3) becomes the second column,

- the column that was stand-by (n°4) now becomes the third column in the series.

This stabilisation process is continued until column n° 2, now the first column, shows
5 a significant pH-drop.

The results of such a carousel process are given in the next table.

Table 4: Continuous carousel process for stabilising maltitol syrup

Columns in service	Syrup processed (l)	pH outlet first column	S-value last column outlet
(1)(2)(3)	1,0	9,3	0,026
	2,0	9,1	0,032
	3,0	8,8	0,042
	4,0	8,2	0,053
(2)(3)(4)	5,0	9,1	0,036
	6,0	9,0	0,044
	7,0	8,7	0,053
	8,0	8,2	0,064
(3)(4)(1)	9,0	9,1	0,038
	10,0	9,0	0,047
	11,0	8,7	0,058
	12,0	8,0	0,067

CLAIMS

1. Process for preparing alkali- and heat-stable sugar alcohol compositions which exhibits an optical density lower than or equal to 0,100 in an S-test, **characterised in that** a sugar alcohol composition is treated with a strong base anion exchange resin in the hydroxide form, at a temperature between 30 °C and 100 °C.
2. Process according to claim 1, **characterised in that** in said process the sugar alcohol composition is fed to a column-system containing a strong base anion exchange resin in the hydroxide form with a volume throughput of ≤ 6 bed volumes (BV)/hour.
3. Process according to claim 2, **characterised in that** a single column-system is used.
4. Process according to claim 2, **characterised in that** a multiple column-system is used, in which at least part of the columns of the system is used in a regeneration mode, while the remaining columns are used in a service mode, comprising the steps of stabilisation and simultaneous decolourisation.
5. Process according to any one of claims 2 to 4, **characterised in that** the volume throughput is between 0,1 and 1 BV/hour.
6. Process according to claim 5, **characterised in that** the volume throughput is between 0,2 and 0,8 BV/hour.
7. Process according to any one of claims 1 to 6, **characterised in that** said sugar alcohol composition has a conductivity value less than 100 $\mu\text{S}/\text{cm}$ before treatment with the strong base anion exchange resin.

8. Process according to claim 7, **characterised in that** said sugar alcohol composition has a conductivity less than 50 $\mu\text{S}/\text{cm}$ before treatment with the strong base anion exchange resin.
- 5
9. Process according to any one of claims 1 to 8, **characterised in that** said strong base anion exchange resin belongs to the thermally stable-type category.
10. Process according to any one of claims 1 to 8, **characterised in that** said strong base anion exchange resin is of the styrenic type I, type II or type III.
- 10
11. Process according to any of claims 1 to 8, **characterised in that** said strong base anion exchange resin is of the acrylic resin type.
- 15
12. Process according to claim 10 or 11, **characterised in that** when using a styrenic type I or type III, or an acrylic type resin, a column temperature is used between 45 °C and 70 °C.
13. Process according to claim 11, **characterised in that** when using a styrenic type II resin, a column temperature is used which is less than 45 °C.
- 20
14. Process according to claim 10, **characterised in that** when using a thermally stable resin, a column temperature is used which is more than 75 °C.
- 25
15. Process according to any one of claims 1 to 14, **characterised in that** said sugar alcohol composition is prepared by hydrogenating a starch hydrolysate, obtained from an acid conversion, a combined acid-enzymatic conversion or a multiple enzyme conversion of starch.
- 30
16. Process according to any one of claims 1 to 14, **characterised in that** said sugar alcohol composition is prepared by hydrogenating reducing sugars belonging to

the categories of keto- or aldopentoses, keto- or aldohexoses, disaccharides or non-starch oligosaccharide mixtures.

- 5 17. Process according to any one of the preceding claims, **characterised in that** said sugar alcohol composition has a pH-value between 8,5 and 9,5 when sorting from the strong base anion exchange resin.
18. Sorbitol composition containing at least 95% sorbitol on dry substance and exhibiting an optical density lower than 0,02 in an S-test.
- 10 19. Sorbitol composition according to claim 18, characterised in that it contains at least 99 % sorbitol on dry substance and exhibits an optical density of lower than 0,01 in an S-test.

INTERNATIONAL SEARCH REPORT

International Application No

PCT/EP 02/14916

A. CLASSIFICATION OF SUBJECT MATTER
 IPC 7 C07C29/74 C07C31/26

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
 IPC 7 C07C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)
 EPO-Internal, WPI Data, PAJ

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category °	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	EP 0 711 743 A (ROQUETTE FRERES) 15 May 1996 (1996-05-15) cited in the application column 8, line 36 - line 54; example 1	1-17
A	EP 1 095 925 A (ROQUETTE FRERES) 2 May 2001 (2001-05-02) paragraph '0020!	18,19

Further documents are listed in the continuation of box C.

Patent family members are listed in annex.

° Special categories of cited documents :

- *A* document defining the general state of the art which is not considered to be of particular relevance
- *E* earlier document but published on or after the international filing date
- *L* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- *O* document referring to an oral disclosure, use, exhibition or other means
- *P* document published prior to the international filing date but later than the priority date claimed

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- *X* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- *Y* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- * & * document member of the same patent family

Date of the actual completion of the international search

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INTERNATIONAL SEARCH REPORT

International application No.
PCT/EP 02/14916

Box I Observations where certain claims were found unsearchable (Continuation of item 1 of first sheet)

This International Search Report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. Claims Nos.:
because they relate to subject matter not required to be searched by this Authority, namely:

2. Claims Nos.:
because they relate to parts of the International Application that do not comply with the prescribed requirements to such an extent that no meaningful International Search can be carried out, specifically:

3. Claims Nos.:
because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a).

Box II Observations where unity of invention is lacking (Continuation of item 2 of first sheet)

This International Searching Authority found multiple inventions in this international application, as follows:

see additional sheet

1. As all required additional search fees were timely paid by the applicant, this International Search Report covers all searchable claims.
2. As all searchable claims could be searched without effort justifying an additional fee, this Authority did not invite payment of any additional fee.
3. As only some of the required additional search fees were timely paid by the applicant, this International Search Report covers only those claims for which fees were paid, specifically claims Nos.:
4. No required additional search fees were timely paid by the applicant. Consequently, this International Search Report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

Remark on Protest

- The additional search fees were accompanied by the applicant's protest.
- No protest accompanied the payment of additional search fees.

FURTHER INFORMATION CONTINUED FROM PCT/ISA/ 210

This International Searching Authority found multiple (groups of) inventions in this international application, as follows:

1. claims: 1-17

Process for the preparation of sugar alcohol compositions

2. claims: 18,19

Sorbitol compositions

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT/EP 02/14916

Patent document cited in search report	Publication date	Patent family member(s)	Publication date				
EP 0711743	A	15-05-1996	FR 2726829 A1	15-05-1996			
			AT 181723 T	15-07-1999			
			AU 700702 B2	14-01-1999			
			AU 3772795 A	16-05-1996			
			BR 9505131 A	11-03-1997			
			CA 2162527 A1	11-05-1996			
			CN 1130502 A , B	11-09-1996			
			DE 69510520 D1	05-08-1999			
			DE 69510520 T2	05-01-2000			
			EP 0711743 A1	15-05-1996			
			ES 2133692 T3	16-09-1999			
			FI 955399 A	11-05-1996			
			HU 74979 A2	28-03-1997			
			JP 8228719 A	10-09-1996			
			PL 311337 A1	13-05-1996			
			US 5773604 A	30-06-1998			
			ZA 9509305 A	04-11-1996			
			EP 1095925	A	02-05-2001	FR 2800370 A1	04-05-2001
						AT 247615 T	15-09-2003
AU 6960700 A	03-05-2001						
BR 0007264 A	16-10-2001						
CA 2324117 A1	28-04-2001						
CN 1327439 T	19-12-2001						
DE 60004620 D1	25-09-2003						
EP 1095925 A1	02-05-2001						
WO 0130731 A1	03-05-2001						
HU 0105095 A2	29-04-2002						
JP 2003512445 T	02-04-2003						
NO 20012912 A	13-06-2001						
PL 348525 A1	03-06-2002						
US 6417346 B1	09-07-2002						
ZA 200005942 A	20-04-2001						