

PATENT COOPERATION TREATY

PCT

INTERNATIONAL PRELIMINARY EXAMINATION REPORT
(PCT Article 36 and Rule 70)

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Applicant's or agent's file reference 1949-19 wo cvp	FOR FURTHER ACTION See Notification of Transmittal of International Preliminary Examination Report (Form PCT/PEA/416)	
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Applicant AMYLUM EUROPE NV		

- This international preliminary examination report has been prepared by this International Preliminary Examining Authority and is transmitted to the applicant according to Article 36.
- This REPORT consists of a total of 6 sheets, including this cover sheet.
 - This report is also accompanied by ANNEXES, i.e. sheets of the description, claims and/or drawings which have been amended and are the basis for this report and/or sheets containing rectifications made before this Authority (see Rule 70.16 and Section 607 of the Administrative Instructions under the PCT).

These annexes consist of a total of 17 sheets.

- This report contains indications relating to the following items:
 - I Basis of the opinion
 - II Priority
 - III Non-establishment of opinion with regard to novelty, inventive step and industrial applicability
 - IV Lack of unity of invention
 - V Reasoned statement under Rule 66.2(a)(ii) with regard to novelty, inventive step or industrial applicability; citations and explanations supporting such statement
 - VI Certain documents cited
 - VII Certain defects in the international application
 - VIII Certain observations on the international application

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**INTERNATIONAL PRELIMINARY
EXAMINATION REPORT**

International application No. PCT/EP 02/14916

I. Basis of the report

1. With regard to the **elements** of the international application (*Replacement sheets which have been furnished to the receiving Office in response to an invitation under Article 14 are referred to in this report as "originally filed" and are not annexed to this report since they do not contain amendments (Rules 70.16 and 70.17)*):

Description, Pages

1-14 filed with telefax on 29.11.2004

Claims, Numbers

1-19 filed with telefax on 29.11.2004

2. With regard to the **language**, all the elements marked above were available or furnished to this Authority in the language in which the international application was filed, unless otherwise indicated under this item.

These elements were available or furnished to this Authority in the following language: , which is:

- the language of a translation furnished for the purposes of the international search (under Rule 23.1(b)).
- the language of publication of the international application (under Rule 48.3(b)).
- the language of a translation furnished for the purposes of international preliminary examination (under Rule 55.2 and/or 55.3).

3. With regard to any **nucleotide and/or amino acid sequence** disclosed in the international application, the international preliminary examination was carried out on the basis of the sequence listing:

- contained in the international application in written form.
- filed together with the international application in computer readable form.
- furnished subsequently to this Authority in written form.
- furnished subsequently to this Authority in computer readable form.
- The statement that the subsequently furnished written sequence listing does not go beyond the disclosure in the international application as filed has been furnished.
- The statement that the information recorded in computer readable form is identical to the written sequence listing has been furnished.

4. The amendments have resulted in the cancellation of:

- the description, pages:
- the claims, Nos.:
- the drawings, sheets:

5. This report has been established as if (some of) the amendments had not been made, since they have been considered to go beyond the disclosure as filed (Rule 70.2(c)).

(Any replacement sheet containing such amendments must be referred to under item 1 and annexed to this report.)

6. Additional observations, if necessary:

**INTERNATIONAL PRELIMINARY
EXAMINATION REPORT**

International application No. **PCT/EP 02/14916**

**V. Reasoned statement under Article 35(2) with regard to novelty, inventive step or industrial applicability;
citations and explanations supporting such statement**

1. Statement

Novelty (N)	Yes: Claims	1-17
	No: Claims	18,19
Inventive step (IS)	Yes: Claims	1-17
	No: Claims	18,19
Industrial applicability (IA)	Yes: Claims	1-19
	No: Claims	

2. Citations and explanations

see separate sheet

Re Item V

Reasoned statement under Article 35(2) with regard to novelty, inventive step or industrial applicability; citations and explanations supporting such statement

Reference is made to the following document:

D1: EP 0 711 743 A (Roquette Frères) 15 May 1996

D2: Fluka Chemika-BioChemika Catalogue 1995/96, pages 1382-1383

The document D2 was not cited in the International Search Report.

1. Subject-matter

The present application concerns a process for the preparation of alkali- and heat- stable sugar alcohol compositions with an optical density lower than or equal to 0.10 in an S-test as defined in the application, in which the crude sugar alcohol is treated with a strongly basic anion exchange resin in the hydroxide form at a temperature between 30 °C and 100 °C. The application also concerns a sorbitol composition containing at least 95% sorbitol and having an optical density of less than 0.02 in an S-test.

2. Novelty

The present application does not meet the criteria of Article 33(1) PCT, because the subject-matter of claims 18,19 is not new in the sense of Article 33(2) PCT.

- 2.1 Document D1 describes polyol compositions having an optical density less than or equal to 0.100 in an S-test (claim 1). These compositions are prepared first treating the polyol composition with glucose oxidase at pH 5.0 at 35 °C, and then passing it through a strong cation exchange resin (IR200 C) and then a strong anion exchange resin (IRA 900) (example 1) at an undisclosed temperature.

Thus, the process of the present application differs from the prior art process in D1 in the nature of the ion exchange resin employed. Consequently, the subject-matter of the present claims 1-17 is new and the requirements of Article 33(1) PCT appear to be met.

- 2.2 Claims 18,19 concern a sorbitol composition "prepared by the process according to any of the preceding claims". A claim defining a product in terms of a process is construed

as a claim to the product per se that possesses the characteristics derived from the process (PCT International Search and Preliminary Examination Guidelines, paragraph A5.26), i.e. the product is **obtainable** by the process. Thus, **a product is not rendered new merely by the fact that it is produced by a new process.**

The sorbitol composition of claims 18,19 contains at least 95 % sorbitol (99 % in claim 19) and has an optical density lower than 0.02 (claim 18) or 0.01 (claim 19) in an S-test. According to the present application, the S-test is a measure of the level impurities (such as residual carbohydrate, presumably) in the sorbitol which cause discolouration on heat treatment in the presence of sodium hydrogencarbonate and ammonia (page 5, lines 16-31).

Document D2 discloses (product code 85529) D-sorbitol which is greater than 99.5 % pure by HPLC analysis and forms a colourless 1M aqueous solution. Since, as a rule, conventional methods for the purification of low molecular weight organic compounds are within the common knowledge of the skilled person, a document disclosing low molecular weight compound normally makes this compound available to the public in all grades of purity. Although optical density of the sorbitol of D2 in an S-test are not known, there is nothing in D2 to suggest that the sorbitol composition of the present claims 18,19 form an exception to this general rule.

Thus, the subject-matter of claims 18,19 does not appear to be new and the requirements of Article 33(2) PCT are not met.

3. Inventive step

- 3.1 Document D1 is regarded as being the closest prior art to the subject-matter of claims 1-17, and discloses a process for the preparation of a polyol composition which includes the step of passing it through a series of ion exchange resins as described in paragraph 2 above. The subject-matter of claim 1 of the present application differs from this known process in the nature of the ion exchange resin employed, namely only a strongly basic anion exchange resin in the hydroxide form, and possibly in the temperature at which the process is carried out.

Table 2 of the present application indicates that the procedure of the present application can reduce the optical density of a sorbitol composition from 1.11 to 0.042 and even as

low as 0.014. Similar reductions in optical density in the use of the same process on other sugar alcohol compositions are presented in tables 3 and 4. The prior art process of D1 (example 1) including the glucose oxidase treatment step reduces the optical density of a sorbitol composition from 0.600 only as far as close to 0.070.

Consequently, the problem being solved by the present invention is the provision of an improved process for the decolouration of sugar alcohol compositions. The solution proposed in claims 1-17, namely the use of a strongly basic anion exchange resin in the hydroxide form, can be considered to involve an inventive step (Article 33(1) PCT in the sense of Article 33(3) PCT) since it is not obvious from D1, or any where else in the prior art, that the use of these ion exchange resins would lead to a larger improvement in the optical density of the sugar alcohol compositions.

- 3.2 In the absence of novelty (see paragraph 2.2 above), no inventive step can be acknowledged for the subject-matter of claims 18,19.

Amended description under art. 34(2)b PCT

**PROCESS FOR PREPARING ALKALI- AND HEAT-STABLE SUGAR
ALCOHOL COMPOSITIONS AND A SORBITOL COMPOSITION**

5

The invention relates on the one hand to a process for preparing alkali- and heat-stable sugar alcohol compositions which exhibits an optical density lower than or equal to 0,100 in an S-test. On the other hand, the invention relates to sorbitol compositions prepared by such a process.

10

Alkali- and heat-stability of sugar-alcohols is important in all those applications where colour formation under these conditions must be prohibited. This is the case, e.g. where polyol compositions are used as humectants in tooth-pastes containing alkaline abrasives, as building blocks of polyether polyols, or as starters for preparing sorbitan esters. Colouring of end-products containing these polyols is often due to the presence of colour-forming precursors, including residual reducing sugars, in the sugar-alcohol compositions used.

15

This problem is quite well known and a number of solutions have already been proposed to improve alkali- and heat-stability of such polyol compositions.

20

In JP 63079844, a method is described in which an aqueous sugar alcohol solution is adjusted to a pH-value of between 8 and 13, followed by a (discontinuous) heating step at temperatures varying between 90°C and 220°C. The resulting product is then purified by passing the polyol solution through a strongly acidic cation exchange resin, a strong base anion exchange resin and a mixed bed resin.

25

In EP 0 711 743, a similar process is disclosed in which the polyol composition is first stabilised by means of an oxidation, a fermentation or a caramelisation step, followed by a purification of the solution. The purification step is comparable to the one disclosed in JP 63079844.

30

In EP 1 095 925, a purification process is disclosed comprising a first treatment on a strong acid cationic exchange resin at a temperature below 50°C, preferably below 40°C, followed by a treatment on a strong basic anionic resin and a mixed bed resin.

5 Methods for removing aldehydes and other reactive impurities, and/or stabilising colour in glycerol and glycol aqueous solutions have been discussed in FR 1 546 472 and US 6,187,973 respectively. In both cases, a strong base anion exchange resin in the hydroxide form was converted into the bisulphite form, in order to treat the aqueous polyol solutions.

10

The major disadvantage of the above-cited processes resides in their complexity. Indeed, first a chemical stabilisation step is needed, followed by a quite complicated purification step. During this stabilisation step, high pH-values are used. In addition, it is necessary to use quite high temperatures (> 90°C) and long reaction
15 times to obtain the necessary stabilisation.(see JP63079844 and EP711743). This results in a rather important chemicals consumption during the chemical reaction, and later on, for the regeneration of the different ion exchange resins.

In addition, two separate steps are needed to arrive at the desired result. The
20 equipment needed, therefore comprises a reactor to perform the stabilisation step, and at least two ion exchange resin batteries to perform the purification step.

In the case of the bisulphite-type resins, the use thereof proves to be inefficient when higher polyols such as pentitols, hexitols and/or hydrogenated starch hydrolysates are
25 treated in the expectation of obtaining alkali- and heat stable products.

The purpose of the invention is to provide a simple process for preparing alkali- and heat-stable sugar alcohol compositions which exhibit an optical density lower than or equal to 0,100 in an S-test. This process results in a reduced consumption of
30 chemicals and provides a combined one step alkaline stabilisation and

decolourisation process. This one step process can be operated in a continuous matter.

5 This object is obtained by providing a process for preparing alkali- and heat-stable sugar alcohol compositions which exhibits an optical density lower than or equal to 0,100 in an S-test, in which a sugar alcohol composition is treated with a strong base anion exchange resin in the hydroxide form, at a temperature between 30 °C and 100 °C.

10 In a preferred process according to the invention, the sugar alcohol composition is fed to a column-system containing a strong base anion exchange resin in the hydroxide form with a volume throughput of ≤ 6 bed volumes (BV)/hour.

15 "Bed volume" is hereby defined as the total volume of resin used during the stabilisation step, be it in a one column- or in a multiple column-system.

20 When a multiple column-system is used, at least part of the columns of the system is used in a regeneration mode, while the remaining columns are used in a service mode, comprising the steps of stabilisation and simultaneous decolourisation.

In a more preferred process according to the invention, the volume throughput is between 0,1 and 1 BV/hour.

25 Most preferably, the volume throughput is between 0,2 and 0,8 BV/hour.

Before treatment with the strong base anion exchange resin, said sugar alcohol composition has preferably a conductivity value less than 100 $\mu\text{S}/\text{cm}$, more preferably 50 $\mu\text{S}/\text{cm}$.

30 The strong base anion exchange resin preferably belongs to one of the categories:
- the thermally stable-type category;

- the styrenic type I, type II or type III; or
- the acrylic resin type.

5 When using a styrenic type I or type III, or an acrylic type resin, a column temperature is preferably used between 45 °C and 70 °C.

When using a styrenic type II resin, a column temperature is preferably used less than 45 °C.

- 10 When using a thermally stable resin, a column temperature is preferably used which is more than 75 °C.

15 On the one hand, said sugar alcohol composition can be prepared by hydrogenating a starch hydrolysate, obtained from an acid conversion, a combined acid-enzymatic conversion or a multiple enzyme conversion of starch.

20 On the other hand, said sugar alcohol composition can be prepared by hydrogenating reducing sugars belonging to the categories of keto- or aldopentoses, keto- or aldohexoses, disaccharides or non-starch oligosaccharide mixtures.

In a preferred process according to the invention, said sugar alcohol composition has a pH-value between 8,5 and 9,5 when sorting from the strong base anion exchange resin.

- 25 The purpose of the invention is furthermore to provide a sorbitol composition prepared by a process as described above, which has a great alkali- and heat stability.

30 This purpose is obtained by providing a sorbitol composition prepared by a process according to the invention, which contains at least 95% sorbitol on dry substance and which exhibits an optical density lower than 0,02, more preferably containing at least 99 % sorbitol on dry substance and exhibiting an optical density of lower than 0,01.

These particular and unexpected sorbitol compositions can be the result of the process of the invention. This doesn't exclude the fact that also other processes can be used for obtaining sorbitol compositions with the characteristics as mentioned
5 above.

To prepare alkali- and heat-stable sugar alcohol compositions which exhibits an optical density lower than or equal to 0,100 in an S-test, which is within the terms of EP 0 711 743, a one step process is used in which all the colour precursors are
10 eliminated, and the resulting sugar alcohol syrup become colourless and alkali- and heat-stable. Thereby surprisingly low reaction temperatures are used, i.e. between 30 and 100°C. These temperatures are sufficient to provide the desired result, as expressed by the S-value, when a strong base anion exchange resin in the hydroxyl form is used to treat the sugar alcohol compositions.

15

The S-test relies on a spectrophotometric measurement applied to the products to be tested.

As used herein, the "S-test" is the following test procedure:

- the aqueous polyol syrup to be tested is brought to a solids content of
20 40 % by weight, if needed by concentration or by aqueous dilution;
- to 5 ml of this solution are added 500 mf of sodium hydrogencarbonate of ultrapure quality, sold, for example, under the name of RP NormapurTM, analytical grade, by the company Prolabo, 65 Bd Richard Lenoir, Paris, France, and 250 mg of an aqueous solution
25 containing 20 % of ammonia;
- the whole is mixed and heated for 2 hours on a steam bath at 100 °C without stirring being applied;
- the solution is brought to 20 °C and the optical density of the solution thus
30 obtained is measured at a wavelength of 420 nm by virtue of a spectrophotometer such as that marketed by Perkin-Elmer under the trademark Lambda 5 UV/VIS Spectrophotometer.

In a process according to the invention, a single or multiple column-system is used, containing a strong base anion exchange resin in the hydroxide form with a volume throughput of ≤ 6 bed volumes (BV)/hour.

5

An advantageous continuous multiple column-system for this invention is known as an ISEP- or as a CSEP-configuration. Thereby, a part of the columns in the system is used in the regeneration mode, while the remaining columns are used for the stabilisation and simultaneous decolourisation of the substrate, i.e. a service mode. Thereby, a column operating in the service mode will become deactivated, after that a certain quantity of substrate is processed by that column. This deactivation can be observed by monitoring the pH of the syrup leaving that column. This "exhausted" column is then switched to the regeneration mode and replaced by a regenerated column. The volume throughput is preferably between 0,1 and 1 BV/hour, more preferably between 0,2 and 0,8 BV/h.

10
15

The strong base anion exchange resins used in this invention belong to the styrenic type I, type II, or type III categories, to the acrylic resin type, and to the thermally stable-type categories.

20 The styrenic type I-category comprises resin types such as Amberlite IRA404, FPA90 and Amberjet 4400, Dowex Marathon 11 and Lewatit M500.

A typical representative of the type II-resins is the styrenic resin type Dowex 22.

A typical representative of the type III-resins is the styrenic resin type Purolite A555.

The acrylic resin type categorie is represented e.g. by Amberlite IRA458 and
25 Amberlite FPA98.

In the category of the thermally stable strong base anion resins, Diaion TSA1200 is a typical example.

When using styrenic type I or type III resins, or acrylic type resins, column
30 temperature is preferably between 45°C and 70°C.

When using styrenic type II resins, column temperature is preferably $< 45^\circ\text{C}$.

When using thermally stable resins, column temperature is preferably $>75^{\circ}\text{C}$.

The sugar alcohol composition is obtained via the hydrogenation of reducing sugar compositions. Typical reducing sugar compositions are starch hydrolysates, but also
5 other reducing sugars including keto- and aldopentoses, keto- and aldohexoses, disaccharides (e.g. lactose, maltose, isomaltose, isomaltulose) and non-starch oligosaccharide mixtures are covered by this term. The term starch hydrolysates refers to those compositions obtained via an acid conversion, a combined acid-enzymatic conversion, or a multiple enzyme conversion of starch, including chemical
10 or enzymatic isomerisation. Typical representatives here are maltodextrins, "standard" glucose syrups, maltose syrups, high DE conversion syrups such as 96DE and 99DE glucose syrups, isoglucoses and crystallisation mother liquors.

The starch may be of cereal, tuber root or leguminous origin.

15 The sugar alcohol substrate needing a stabilisation treatment, preferably has a conductivity value $< 100 \text{ S/cm}$, more preferable $< 50 \mu\text{S/cm}$, before treatment with the strong base anion exchange resin. Such sugar alcohol compositions can be obtained when using a noble metal hydrogenation catalyst, for example Ru, Pt or Pd; or by first removing dissolved residual metal ions from the substrate, in those cases
20 where transition metals such as for example Ni, Co, Cu or Fe are used as the catalysts.

The processed, colour-stable sugar alcohols thus obtained typically have a pH-value varying between 8,5 and 9,5 when sorting from the strong base anion exchange resin.

25 These sugar alcohol compositions can then be used as such or further processed by means of a mixed bed resin or a weak acid cation exchange resin, thereby providing a syrup having a $\text{pH} = 4 - 7$, preferably $\text{pH} = 5 - 6,5$.

Sorbitol compositions with great alkali- and heat stability can be obtained by the
30 process of the invention as described above. This doesn't exclude the fact that also

other processes can be used for obtaining such particular and unexpected sorbitol compositions.

5 The sorbitol composition comprise at least 95% sorbitol on dry substance and exhibits an optical density lower than 0,02.

10 In a more preferred embodiment of the invention, the sorbitol composition comprise at least 99 % sorbitol on dry substance and exhibits an optical density of lower than 0,01.

In order to illustrate this invention, a number of examples are provided hereunder.

EXAMPLES:

15 Comparative example 1:

20 A sorbitol syrup containing 96% sorbitol and showing an S-value of 1,10 before stabilisation is submitted to an alkaline heat treatment. Thereby the pH is brought to pH = 11,2 with 1N NaOH and the syrup is heated for 2 hours at different temperatures, followed by a refining step. After refining the stability of the thus processed syrup is measured by means of the S-test. Thereby it is clearly illustrated that sufficiently high temperatures are needed to obtain sufficient stabilisation within a reasonable time. The effect of the stabilisation conditions is clearly illustrated in the following table.

25

Stabilisation conditions	S-value
Unstabilised	1,10
2h at 80°C and pH=11,2	0,35(is>0,1)
2h at 100°C at pH=11,2	0,021

Comparative example 2:

In this example it is shown that the use of strong base anionic exchange resins, in the sulphite form, is not suitable to obtain alkali- and heat-stable polyols, comparable
5 with the polyols disclosed in EP 0 711 743.

A strong base anion exchange resin in the hydroxide form was thereby converted to the bisulphite form by passing a 2 mol/l solution of NaHSO₃ through a bed of the resin until the bisulphite inlet and outlet concentrations were essentially the same. This resin was then washed with 10 volumes of water per volume of resin.

10 A unrefined sorbitol solution, having a S-value of 0,85 (as determined by the S-test of EP711743) was then passed through the bisulphite-treated resin at ambient temperature.

The treated solutions did not fulfil the requirements set forward by EP 0 711 743, in order to claim alkali/heat stability for these polyols, as illustrated in table 1.

15

Table 1: Treatment of sorbitol by anion exchange resin in bisulphite form:

Sorbitol(99,5%)	S-value
Untreated	0,85
Dowex 22 (1)	0,72
Lewatit M500 (1)	0,64

(1): Dowex 22(type II) and Lewatit M500 (type I) are both strong base anion
20 exchange resins, transformed in the bisulphite form as described above.

Example 1:

Four different columns are filled with 100 ml of a different strong base anionic exchange resin in the hydroxide form.

- 5 The resins used are: Amberlite FPA 90, Amberlite IRA 458, Amberjet 4400 and Purolite A555.

The operating temperature for the columns containing Amberlite FPA90 and Amberjet 4400 is 60°C. The operating temperature for the columns containing Amberlite IRA458 and Purolite A555 is 55°C.

- 10 The unrefined sorbitol syrup used is obtained by the hydrogenation of a 96 DE-glucose syrup and has a concentration of about 50% d.s.. The S-value of the syrup is 1,11.

The syrup is passed through the columns at 0,33 BV/hour. The S-value of the processed substrate is determined in the mixture of the collected volumes of the first

- 15 25 bed volumes.

Table 2: Sorbitol syrup stabilisation via strong base anionic resin treatment as expressed by their S-value:

Sorbitol syrup substrate	S-value
untreated	1,11
Amberlite FPA 90	0,042
Amberlite IRA 458	0,016
Amberjet 4400	0,016
Purolite A555	0,014

Example 2:

The same procedure as in example 1 was followed. The substrate treated now is sorbitol syrup, obtained by the hydrogenation of a D99 dextrose syrup. The sorbitol
5 syrup has a 99,4% sorbitol content and an S-value of 0,87 before stabilisation.

This syrup is again passed through the columns at 0,33 BV/hour. The processed sorbitol syrup is collected. A mixed sample corresponding to the mixture of the first 30 bed volumes is submitted to the S-test.

10 **Table 3: D99 Sorbitol syrup stabilisation via strong base anionic resin treatment as expressed by their S-value:**

D99 Sorbitol syrup substrate	S-value
untreated	0,87
Amberlite FPA 90	0,007
Lewatit M500	0,006
Purolite A555	0,007

15 Example 3:

The same procedure as in example 1 was followed. The substrate treated now is a maltitol syrup, obtained by the hydrogenation of a high maltose syrup containing
20 about 65 % maltose, 8 % glucose and 20 % maltotriose. The maltitol syrup has an S-value of 2 before refining.

This syrup is again passed through the columns at 0,33 BV/hour. The processed maltitol syrup is collected. A mixed sample corresponding to the mixture of the first 25 bed volumes is submitted to the S-test.

25

Table 4: Maltitol syrup stabilisation via strong base anionic resin treatment as expressed by their S-value:

Maltitol syrup substrate	S-value
Untreated	2,0
Amberlite FPA 90	0,065
Amberlite IRA 458	0,08
Amberjet 4400	0,11
Purolite A555	0,09

- 5 This example shows that the substrate composition, the original S-value of the substrate, as well as the ion exchange resin used, may affect the refining capacity.

Example 4:

10

One column is filled with 100 ml of a strong base thermally stable anion exchange resin (Diaion TSA1200). The substrate treated is a maltitol syrup, obtained by the hydrogenation of a high maltose syrup containing about 65% maltose, 8% glucose and 20% maltotriose. The maltitol syrup has an S-value of 2,2 before refining.

- 15 The syrup (50% d.s.) is passed, at 90°C, through the column at a speed of 2BV/hour. The evolution of the S-value as a function of the number of bed volumes processed is given in the next table.

BV treated	S-value
5	0,012
10	0,013
20	0,022
30	0,028
40	0,037

Example 5:

One column is filled with 100 ml of a strong base styrenic type II anion exchange resin (Dowex 22). The substrate treated is a sorbitol syrup (96 % sorbitol), obtained by the hydrogenation of a 96DE glucose syrup. The sorbitol syrup has an S-value of 0,81 before stabilisation.

The syrup (50 % d.s.) is passed, at 40°C, through the column at a speed of 0,2 BV/hour. After 24 hour the S-value of the collected volume was determined. The value of the collect volume was $S = 0,036$.

10

Example 6:

In this example a multiple column system is used to refine a maltitol syrup having a S-value of 2,2 before treatment. The maltitol syrup is obtained by the hydrogenation of a high maltose syrup containing about 65 % maltose, 8 % sorbitol and 20 % maltotriose.

Four columns are each filled with 100 ml Amberlite IRA458 resin in the hydroxyl form. Three out of four columns are connected in series. The maltitol syrup is passed through these three columns at a speed of 100 ml/h. (0,33 BV/h - 300 ml resin in use). The temperature of the columns is 55°C. The treated syrup is collected in fractions of 1 liter, and of each such fraction the S-value is determined. In a spot sample, at the outlet of the first column, the pH is measured on the moment that a 1 liter fraction is collected at the outlet of the third (last) column of the stabilisation cycle. When the pH shows a significant drop compared to the pH of the previous sampling, then:

- this first column (n°1) is taken out of service and is regenerated
- the second column (n°2) now becomes the first column, and receives the substrate to be treated,
- the third column (n°3) becomes the second column,

30

- the column that was stand-by (n°4) now becomes the third column in the series.

This stabilisation process is continued until column n° 2, now the first column, shows a significant pH-drop.

The results of such a carousel process are given in the next table.

Table 5: Continuous carousel process for stabilising maltitol syrup

Columns in service	Syrup processed (l)	pH outlet first column	S-value last column outlet
(1)(2)(3)	1,0	9,3	0,026
	2,0	9,1	0,032
	3,0	8,8	0,042
	4,0	8,2	0,053
(2)(3)(4)	5,0	9,1	0,036
	6,0	9,0	0,044
	7,0	8,7	0,053
	8,0	8,2	0,064
(3)(4)(1)	9,0	9,1	0,038
	10,0	9,0	0,047
	11,0	8,7	0,058
	12,0	8,0	0,067

Amended claims under art. 34(2)b PCT

1. Process for preparing alkali- and heat-stable sugar alcohol compositions which exhibits an optical density lower than or equal to 0,100 in an S-test, **characterised in that** a sugar alcohol composition is treated with a strong base anion exchange resin in the hydroxide form, at a temperature between 30 °C and 100 °C.
2. Process according to claim 1, **characterised in that** in said process the sugar alcohol composition is fed to a column-system containing a strong base anion exchange resin in the hydroxide form with a volume throughput of ≤ 6 bed volumes (BV)/hour.
3. Process according to claim 2, **characterised in that** a single column-system is used.
4. Process according to claim 2, **characterised in that** a multiple column-system is used, in which at least part of the columns of the system is used in a regeneration mode, while the remaining columns are used in a service mode, comprising the steps of stabilisation and simultaneous decolourisation.
5. Process according to any one of claims 2 to 4, **characterised in that** the volume throughput is between 0,1 and 1 BV/hour.
6. Process according to claim 5, **characterised in that** the volume throughput is between 0,2 and 0,8 BV/hour.
7. Process according to any one of claims 1 to 6, **characterised in that** said sugar alcohol composition has a conductivity value less than 100 $\mu\text{S}/\text{cm}$ before treatment with the strong base anion exchange resin.

8. Process according to claim 7, **characterised in that** said sugar alcohol composition has a conductivity less than 50 $\mu\text{S}/\text{cm}$ before treatment with the strong base anion exchange resin.
- 5
9. Process according to any one of claims 1 to 8, **characterised in that** said strong base anion exchange resin belongs to the thermally stable-type category.
10. Process according to any one of claims 1 to 8, **characterised in that** said strong base anion exchange resin is of the styrenic type I, type II or type III.
- 10
11. Process according to any of claims 1 to 8, **characterised in that** said strong base anion exchange resin is of the acrylic resin type.
- 15
12. Process according to claim 10 or 11, **characterised in that** when using a styrenic type I or type III, or an acrylic type resin, a column temperature is used between 45 °C and 70 °C.
- 20
13. Process according to claim 11, **characterised in that** when using a styrenic type II resin, a column temperature is used which is less than 45 °C.
14. Process according to claim 10, **characterised in that** when using a thermally stable resin, a column temperature is used which is more than 75 °C.
- 25
15. Process according to any one of claims 1 to 14, **characterised in that** said sugar alcohol composition is prepared by hydrogenating a starch hydrolysate, obtained from an acid conversion, a combined acid-enzymatic conversion or a multiple enzyme conversion of starch.
- 30
16. Process according to any one of claims 1 to 14, **characterised in that** said sugar alcohol composition is prepared by hydrogenating reducing sugars belonging to

the categories of keto- or aldopentoses, keto- or aldohexoses, disaccharides or non-starch oligosaccharide mixtures.

5 17. Process according to any one of the preceding claims, **characterised in that** said sugar alcohol composition has a pH-value between 8,5 and 9,5 when sorting from the strong base anion exchange resin.

10 18. Sorbitol composition prepared by a process according to any one of the preceding claims, **characterised in that** it contains at least 95% sorbitol on dry substance and exhibits an optical density lower than 0,02 in an S-test.

19. Sorbitol composition according to claim 18, **characterised in that** it contains at least 99 % sorbitol on dry substance and exhibits an optical density of lower than 0,01 in an S-test.